Name: _____

Block: _____

Hydrates

1. Give the chemical formula and molar mass of iron (II) sulfate heptahydrate.

2. What is the stock name of the compound $CoSO_4 \cdot 7\,H_2O$

3. If 10.0 g of $Na_2CrO_4 \cdot 4\,H_2O$ is heated to constant mass, what will the final mass be? *(Hint: first, find the moles of* $Na_2CrO_4 \cdot 4\,H_2O$*. This will be the same as the moles of anhydrous* Na_2CrO_4 *after heating. Then find the mass of that many moles of anhydrous* Na_2CrO_4*.)*

4. 14.70 g of a hydrate of $CaCl_2$ is heated to dryness. The anhydrous sample has a mass of 11.10 g after evaporating the H_2O. What is the chemical formula of the hydrate?