Electron Configurations

Give the electron configuration (orbital notation—with the arrows) for each of the following elements:

- 1. carbon
- 2. potassium
- 3. silicon
- 4. silver

For each of the following electron configurations, name the element.

5.
$$\frac{11}{1s}$$
 $\frac{11}{2s}$ $\frac{11}{2p}$ $\frac{1}{2p}$

$$6. \ \ \frac{1}{1s} \quad \ \, \frac{1}{2s} \quad \ \, \frac{1}{2p} \quad \frac{1}{2p} \quad \frac{1}{3s} \quad \ \, \frac{1}{3p} \quad \frac{1}{3p} \quad \frac{1}{4s} \quad \quad \frac{1}{1} \quad \frac{1}{3d} \quad \frac{1}{1} \quad \frac{1}{3d}$$

Each of the following electron configurations has something wrong with it. For each one:

- State what the mistake is.
- Re-write the electron configuration correctly, keeping the total number of electrons the same.

7.
$$\frac{1 \downarrow}{1s}$$
 $\frac{1 \downarrow}{2p}$ $\frac{1}{2p}$

8.
$$\frac{1 \downarrow}{1s}$$
 $\frac{1 \downarrow}{2s}$ $\frac{1 \downarrow}{2p}$

9.
$$\frac{1}{1s}$$
 $\frac{1}{2s}$ $\frac{1}{2p}$ $\frac{1}{2p}$ $\frac{1}{3s}$ $\frac{1}{3p}$ $\frac{1}{3p}$ —

$$10. \ \ {\textstyle\frac{11}{1s}} \quad {\textstyle\frac{11}{2s}} \quad {\textstyle\frac{11}{2p}} \ {\textstyle\frac{11}{2p}} \ {\textstyle\frac{11}{3s}} \quad {\textstyle\frac{11}{3p}} \ {\textstyle\frac{11}{3p}} \ {\textstyle\frac{11}{4s}} \quad {\textstyle\frac{1}{1}} \ {\textstyle\frac{1}{4d}} \ {\textstyle\frac{1}{1}} \ -$$